**Colligative Properties In-Class Sheet**

1) Which would have a higher melting point, a 2 m solution of NaBr or a 1 m solution of sugar (which is covalent)? Explain your answer.

2) What would the boiling point of a 0.55 m solution of CuBr2 be? The Kb of water is 0.52 degrees/m.

3) What would the freezing point of the solution in problem 2 be? Kf of water is 1.86 degrees C/m.

4) Explain why more concentrated solutions have lower freezing points than less concentrated solutions.

5) If I had a solution that was simultaneously 0.1 M NaBr *and* 0.1 M sucrose (covalent), what would the boiling point of this solution be?